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Solutions Worksheet 2 Molarity And

Molarity Worksheet # 2
identifiers _____ What
does molarity mean?
Number of moles of
solute. 1 liter solution.

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Molarity And
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What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution? $4.53 \text{ mol LiNO}_3 = 1.59 \text{ M LiNO}_3$
3. 2.85 L soln

Molarity Worksheet 2 ANSWERS - Google Docs

Molarity Problems
Worksheet $M=nV$ $n= \#$
moles V must be in
liters (change if
necessary) 1. What is
the molarity of a 0.30

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Worksheet 1
Molarity And
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liter solution containing
0.50 moles of NaCl? 2.
Calculate the molarity
of 0.289 moles of FeCl₃
dissolved in 120 ml of
solution? 3. If a 0.075
liter solution c...

Molarity and Dilutions Worksheet - Google Docs

6.00 liters of solution?
= 0.500 M NaCl 2.

What is the molarity of
KCl solution containing
1.70 moles of KCl in
3.00 liters of solution?

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3.00 L of Solution 3.

What is the molarity of a solution containing 4.20 moles of sulfuric acid in 300.0 mL of solution? Suppose we want to know the number of moles when given the volume and molarity.

Molar Concentration of Solutions

Solutions Worksheet
#2. (Molarity, Dilutions,
Percent Solutions,
Molality Problems)

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Molarity. Tell how you would prepare a 500. mL of 0.50 M ammonium carbonate solution. Include all necessary equipment and amount of chemical (in grams).

Solutions Worksheet #2 - Georgetown High School

Molarity Problems.
Molarity Problems -
Displaying top 8
worksheets found for
this concept.. Some of

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the worksheets for this concept are Molarity practice problems, Molarity problems work, Work molarity name, Molarity molarity, Molality work 13, Molarity molality osmolality osmolarity work and key, Molarity work w 331, Concentration work w 328.

Molarity Problems Worksheets - Kiddy Math

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#2: Molarity and

Dilution Problems 1)

Describe how you

would prepare 5.00

liters of a 6.00M

solution of potassium

hydroxide. SL 2) How

would you prepare

100.0ml of AM $MgSO_4$

from a stock solution of

2.0 $MgSO_4$? i 00 3) If

1.00l- of water is

added to 3.00 L of a

6.00M solution of what

is the new molarity of

the acid solution?

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Worksheet 2

Chemistry: Molarity of
Solutions Directions:
Solve each of the
following problems.

Show your work and
include units for full
credit. 1. What mass of
the following chemicals
is needed to make the
solutions indicated? a.
1.0 liter of a 1.0 M
mercury (II) chloride
(HgCl_2) solution. b. 2.0
liters of a 1.5 M sodium
nitrate (NaNO_3)

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solution ...

Worksheet 2

Molarity of Solutions - teachnlearnchem.c om

7. How many liters of solution can be produced from 2.5 moles of solute if a 2.0 M solution is needed?
 $2.0 \text{ M} = 2.5 \text{ moles}$
liters of solution
liters of solution = 1.25 L = 1.3 L

8. What would be the concentration of a solution formed when 1.00 g of NaCl are

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Molarity And
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dissolved in water to
make 100.0 mL of
solution? ? mol = 1.00
g NaCl \times 1 mol NaCl
58.5 g ...

Answer Key

**Molarity: Molarity =
1. 2.**

Molarity Worksheet W
331 Everett

Community College
Student Support

Services Program What

is the molarity of the
following solutions

given that: 1) 1.0

moles of potassium

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Answer Key

fluoride is dissolved to make 0.10 L of solution. 2) 1.0 grams of potassium fluoride is dissolved to make 0.10 L of solution.

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We also acknowledge
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Acids and Bases 2 (Worksheet) -

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Key+. 1)++23.5g+of+
NaCl+isdissolvedineno
ughwatertomake.683L
ofsolution. + a)+ What
+is+themolarity)(M)+o
f+the+solution?+++ M
olar+mass+of+NaCl+
=58.44g/mole+
Moles+of+NaCl:+ 23.5
g+NaCl+++1moleNaCl
+++ =++ .402moles+N
aCl+ ++++++
+++++5
8.44gNaCl+ ++ Molarit
y+++ =+++++

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++ moles ++++
++++ = ++++ 0.4
02 moles + NaCl ++++
++ = 0.589 moles + NaCl/
L + = + 0.589 M) NaCl + +
+++++ + + + + + liters
olution 0.683 L of solution
+ + b) ++ How + many +
moles + of + NaCl + are co
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Calculations + for + So lutions + Worksheet + and + Key +

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dissolved to make 0.10
L of solution. 2) 1.0
grams of potassium
fluoride is dissolved to
make 0.10 L of
solution. Molarity
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Solutions Worksheet 2 Molarity And Dilution Problems ...

What is the molarity of an ammonium carbonate solution if the concentration of ammonium ions is 2 M?

What is the concentration of carbonate ions and what is the total concentration of solute particles? $[\text{CO}_3^{2-}] = 1\text{M}$ $[\text{particles}] = 3\text{ M}$. A solution was made by

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Dilution Problems

dissolving 800.0 g of
NaOH in 2.00 L of
water.

Chapter 13 worksheet #1

Answer Key

Water was added to 25
mL of a stock solution
of 5.0 M HBr until the
total volume of the
solution was 2.5 L.

What is the molarity of
the new solution? We
are given the following:

$c_1 = 5.0 \text{ M}$, $V_1 =$
 0.025 L , $V_2 = 2.50 \text{ L}$.

We are asked to find c_2

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Worksheet 2
2, which is the molarity
of the diluted solution.

$$(5.0 \text{ M})(0.025 \text{ L}) = c_2 (2.50 \text{ L})$$

Dilution Problems

Molarity | Introduction to Chemistry

Molarity Problems

Worksheet $M = \frac{n}{V}$ -

$n = \# \text{ moles}$ $V = \text{volume}$ must
be in liters (change if
necessary) - Use M or
mol/L as unit for

molarity 1. What is the
molarity of a 0.30 liter
solution containing

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0.50 moles of NaCl?

Worksheet 2

Molarity Problems

Worksheet - Mrs

Getson's Blog

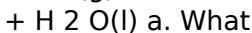
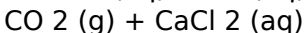
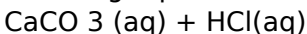
2.2. Calcium carbonate

("limestone") reacts

with hydrochloric acid

according to the

following equation:



a. What mass of calcium

carbonate is needed to

make 1.2 liters of a 1.7

M calcium carbonate

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solution? b. What volume of 3.0 M hydrochloric acid is needed to completely react

Answer Key Molarity and Stoichiometry - Gateway High School

Dilutions Worksheet -
Solutions 1) If 45 mL of water are added to 250 mL of a 0.75 M K_2SO_4 solution, what will the molarity of the diluted solution be?

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Answer Key

$(0.75 \text{ M})(250 \text{ mL}) = M_2(295 \text{ mL})$
 $M_2 = (0.75 \text{ M})(250 \text{ mL}) / (295 \text{ mL}) = 0.64 \text{ M}$

2) If water is added to 175 mL of a 0.45 M KOH solution until the volume is 250 mL, what

Dilutions Worksheet W 329 - Everett Community College

Solutions to Worksheet
#6 1. $2 \times 3.14 \times 4.5 \text{ in}$
 $= 28.26 \text{ in}$ 2. $3.14 \times$
 $18 \text{ in} = 56.52 \text{ in}$

Solutions to Worksheet
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Answer Key

#7 1. $4\text{yd} \times 7\text{yd} = 28$
 yd^2 2. This is a 3-step
problem. Separate the
shape into 2
rectangles, then solve.
Step 1: (Area of 1st
rectangle) $7\text{ft} \times 11\text{ft} =$
 77ft^2 Step 2: (Area ...

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